

KEY

CHEM 3523 001 Fall 2023

Name:

ID:

**Final Exam**

**Answer Sheet:** (The conceptual questions are multiple choice. List the letter that corresponds to the correct answer. For the calculation problems list under **a) the final equation** that gives the solution. Use only symbols, not intermediate numerical results. List under **b) the final numerical result**. Make no mistakes when transferring the answers! Put your names on **both** answer sheets **and** the work pages, and **return all pages!** Conceptual questions are **0.5 pts** each, calculation problems **3+1 pts**. Maximum number of points you can get is **50 pts!**)

**General Infos:**

- No programable calculators, smartphones, smartwatches, tablets, headphones, ... are allowed. Neither any notes or books.
- Any attempt of cheating or other forms of academic dishonesty will result in an automatic "F" for the course.
- **Show a picture ID** when leaving
- Be considerate when you finish early. Consider to stay till end, or at least be quite when leaving earlier to avoid distracting your fellow students!

Conceptual questions:

- 1) c
- 2) b
- 3) a
- 4) b
- 5) a
- 6) c
- 7) a
- 8) b
- 9) a
- 10) a
- 11) b
- 12) b
- 13) d
- 14) a
- 15) a
- 16) b
- 17) a
- 18) a
- 19) a
- 20) a

Problems & Calculations:

- 1a)  $\frac{m_2}{m_1} = \frac{P_2}{P_1} \frac{m_2}{m_1}$
- 1b) ~~0.27~~
- 2a)  $\Delta H_{\text{vap}} = R \cdot \ln \frac{P_2}{P_1} \cdot \frac{T_1 \cdot T_2}{T_2 - T_1}$
- 2b)  $27.02 \frac{\text{kJ}}{\text{mol}}$
- 3a)  $\text{AH}_2$  is oxidized by B
- 3b)  $0.25 \%$
- 4a)  $\alpha = \frac{\lambda}{\lambda_0}$
- 4b)  $0.017$
- 5a)  $E = \frac{h^2}{8m\alpha^2} (1+1+1)$
- 5b)  $1.81 \times 10^{-7} \text{ J}$  ;  $4.52 \times 10^{-10} \text{ m/s}$
- 6a)  $\Delta V_x \approx \frac{h}{2m\Delta x}$
- 6b)  $3.44 \times 10^{-10} \text{ m/s}$
- 7a)  $\Delta E_j \rightarrow j+1 \geq \frac{h^2}{I} (j+1) \lambda \alpha \mu$
- 7b)  $\lambda_D = 2\lambda_H = 1000 \text{ nm}$
- 8a)  $\frac{1}{\lambda} = R_H \left( \frac{1}{j^2} - \frac{1}{i^2} \right)$
- 8b)  $434.2 \text{ nm}$
- 9a)  $E = (n + \frac{1}{2}) h \nu_0$
- 9b)  $1.641 \times 10^{-19} \text{ J}$
- 10a)  $w(2,1,3) = \frac{6!}{2!1!3!}$  ;  $w(1,4,1) = \frac{6!}{1!4!1!}$
- 10b)  $w(2,1,3) = 60$   
 $w(1,4,1) = 30 < 60$